



Grading Period	First Grading Period	Second Grading Period	Third Grading Period	Fourth Grading Period	Fifth Grading Period	Sixth Grading Period
Calendar Dates	Aug. 28 – Oct. 6 (29 days)	Oct.10 – Nov. 10 (24 days)	Nov. 13 – Dec. 15 (20 days)	Jan. 3 – Feb. 16 (32 days)	Feb. 20 - April 13 (33 days)	April 16 – May 31 (33 days)
TEKS	1A, 1B, 1C, 2A, 2B, 2C, 2E, 2G, 2H, 2I, 3A, 3D, 3F, 4A, 4B, 4C, 4D, 5A, 5B, 5C, 6A, 6B, 6C, 6D	1A, 1B, 1C 2A, 2B, 2C, 2D, 2E, 2G, 2F, 2H, 2I, 3A, 3D, 3E, 3F, 5A, 5B, 5C, 6A, 6B, 6C, 6D, 6E, 7A, 7B, 7C, 7D, 7E.	1A, 1B, 1C, 2E, 2F, 2G, 2H, 2I, 3F, 7A, 7B, 8A, 8B, 8C	1A, 1B, 1C, 2C, 2D, 2E, 2F, 2G, 2H, 2I, 3F, 8A, 8D, 8E, 9A, 9B, 9C, 10H	1A, 1B, 1C, 2C, 2D, 2E, 2F, 2G, 2H, 2I, 3F, 10A, 10B, 10C, 10D, 10E, 10F, 9A, 9B, 9C, 10G, 10H, 10I, 10J	1A, 1B, 1C, 2C, 2D, 2E, 2G, 2H, 2I, 3B, 3C, 3D, 3E, 3F, 10G, 10H, 10I, 10J, 11A, 11B, 11C, 11D, 11E, 12A, 12B, 12C
Topic Focus	<p>Unit 01: Laboratory Management: Safety, MSDS, Conservation of Resources</p> <p>Unit 02: Matter: Chemical and Physical Changes, Extensive and Intensive Properties, Classification of Matter</p> <p>Unit 03: Atomic Structure: Development of the Periodic Table, Dalton, Thomson, Rutherford, and Bohr</p>	<p>Unit 03 (Cont'd): Atomic Structure: Development of the Periodic Table, Dalton, Thomson, Rutherford, and Bohr</p> <p>Unit 04: Chemical Bonding: Electron Configuration, Dot Structures, Metallic Properties, VESPR</p> <p>Unit 05: Chemical Formulas: Polyatomic Ions, IUPAC Nomenclature, Acids and Bases, Covalent Compounds</p>	<p>Unit 05 (Cont'd): Chemical Formulas: Polyatomic Ions, IUPAC Nomenclature, Acids and Bases, Covalent Compounds</p> <p>Unit 06 Mole Concept: Moles, Calculation of Ions, Calculate Percent Composition</p>	<p>Unit 07: Chemical Equations: Law of Conservation of Mass, Writing and Balancing Chemical Equations, Acid-Base Reactions</p> <p>Unit 08: Stoichiometry: Stoichiometric Calculations, Reactants and Products, Limiting Reagents, Percent Year</p> <p>Unit 09: Gases: Kinetic Molecular Theory, Boyle's Law</p>	<p>Unit 09 (Cont'd): Gases: Kinetic Molecular Theory, Boyle's Law</p> <p>Unit 10: Solutions: Solubility, Electrolytes, Nonelectrolytes, Concentration of Solutions</p> <p>Unit 11: Acids and Bases: Arrhenius, Bronsted-Lowry, Acid-Base Reactions, Oxidation-Reduction, pH</p>	<p>Unit 11: Acids and Bases: Arrhenius, Bronsted-Lowry, Acid-Base Reactions, Oxidation-Reduction, pH</p> <p>Unit 12: Thermochemistry: Calorimetry, Calculations of Energy Changes, Exothermic, Endothermic</p> <p>Unit 13: Nuclear Chemistry: Alpha, Beta, and Gamma Radiation, Radioactive Decay, Fission, Fusion</p> <p>Unit 14: Chemistry Connections</p>
District Assessment	CBA #1		CBA #2 Semester Exam		CBA #3	Final exams